

Name \_\_\_\_\_

Score \_\_\_\_/50

Date \_\_\_\_\_

## Atomic Structure

Atomic Number = # of Protons = # of Electrons

Mass Number = Protons + Neutrons

Substance	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Sulfide			32			
Chloride					20	
	Mg <sup>2+</sup>				14	
Fluoride			19			
	Na <sup>+</sup>		24			
Phosphide					16	
	K <sup>+</sup>		39			
Bromide					44	
	Zn <sup>2+</sup>		64			
Oxide					8	

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Mass Number = Protons + Neutrons

Substance	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Magnesium			26			
Helium					2	
	Au				122	
Lead			207			
	F		19			
Germanium					42	
	Li		6			
Potassium					20	
	Br		81			
Radon					133	

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## Atomic Structure

1. Using the following words, answer the questions below.

electrons, neutrons, nucleus, protons, shells

- (a) The center of the atom is called the \_\_\_\_\_.
- (b) \_\_\_\_\_ have a positive charge.
- (c) \_\_\_\_\_ have a negative charge.
- (d) \_\_\_\_\_ have no charge.
- (e) \_\_\_\_\_ and \_\_\_\_\_ are contained within the nucleus of the atom.
- (f) Electrons are contained in the \_\_\_\_\_ of an atom.
- (g) The \_\_\_\_\_ and the \_\_\_\_\_ of the atom have the same mass.

2. A neutral atom has 54 protons and 70 neutrons.

- (a) What is the atomic number? \_\_\_\_\_
- (b) What is the mass number of the atom? \_\_\_\_\_
- (c) How many electrons does the atom have? \_\_\_\_\_
- (d) What is the name of the atom? \_\_\_\_\_

3. Assuming all the atoms have neutral charge, complete the following table using the information already provided.

	Protons	Electrons	Neutrons	Mass Number
Atom 1		30	20	
Atom 2			24	50
Atom 3	44			70
Atom 4		12		30
Atom 5	37		43	
Atom 6	19			30